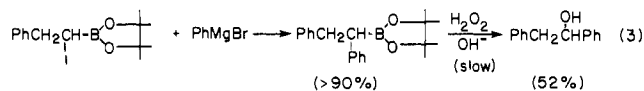


reacted similarly to yield 62% of the boronic ester, which was hydrolyzed with sodium borate in hot aqueous ethanol to the crystalline derivative **7**, mp 104–106 °C (from ether/petroleum ether).⁶ Cyclohexanone with a solution of **2a** gave the enethiol ether **8**, 83% in crude product, 71% after treatment with aqueous ethanolic sodium borate and then sodium hydroxide to remove boron compounds and distillation. The yield of **8** from phenylthiomethyltrimethylsilane was 65%.¹¹ Enethiol ethers were also prepared from **2a** and butyrophenone, bp 130 °C (0.1 mm), 82%,⁶ and from **2a** and benzophenone, mp 69–70 °C,² 71%.

One potential use of the α -(phenylthio)alkaneboronic esters (**3** and **6**) is as precursors of carbonyl compounds. The pinacol boronic ester group has proved unexpectedly resistant to hydrolysis or oxidation, but cleavage of **3a** with *N*-chlorosuccinimide under basic conditions has given high yields of hemithioacetals or dimethyl acetals.¹² Another use is the conversion of **1a** or **3a** (R = C₄H₉, CH₂Ph) to α -iodoalkaneboronic esters (70–77%) with methyl iodide and sodium iodide in dimethylformamide³ for 3 days at 25 °C. α -Haloalkaneboronic esters are of interest for their carbon-carbon bond-forming reactions with Grignard or lithium reagents^{13,14} and as precursors to boronic acids which may bind to enzymes.¹⁵ Carbon-carbon bond formation has been demonstrated with the sequence illustrated (eq 3).



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References and Notes

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Additions and Corrections

Flash Photolysis of Na⁺, C⁻(Ph)₂CH₂CH₂C⁻(Ph)₂, Na⁺. Redox Potential of 1,1-Diphenylethylene and Rate Constant of Dimerization of Its Radical Anion [*J. Am. Chem. Soc.*, **99**, 4612 (1977)]. By H. C. WANG, E. D. LILLIE, S. SLOMKOWSKI, G. LEVIN, and M. SZWARC,* Department of Chemistry, State University of New York, College of Environmental Science and Forestry, Syracuse, New York 13210.

Because of an inadvertent error, the "wrong" Figure 7 was published. The correct figure is presented here.

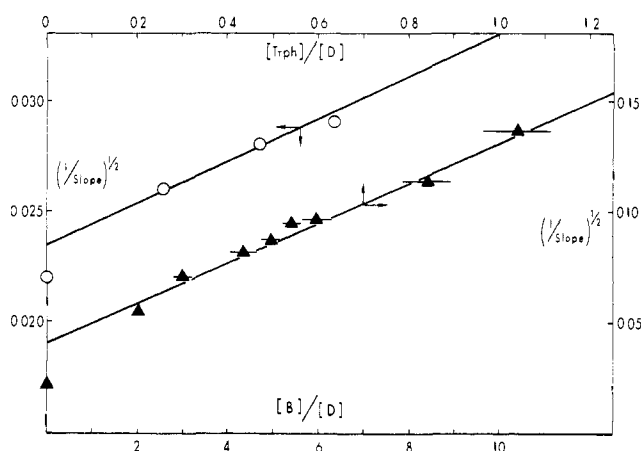


Figure 7. The square root of reciprocals of the corrected slopes of the lines $1/\Delta(\text{OD } 470)$ vs. time (see caption to Figure 6) plotted as functions of $[\text{Trph}]/[\text{D}]$ (triangles) or $[\text{B}]/[\text{D}]$ (circles).

Cyclic Peptides. 17. Metal and Amino Acid Complexes of *cyclo*(Pro-Gly)₄ and Analogues Studies by Nuclear Magnetic Resonance and Circular Dichroism [*J. Am. Chem. Soc.*, **99**, 4788 (1977)]. By VINCENT MADISON, CHARLES M. DEBER, and ELKAN R. BLOUT,* Department of Biological Chemistry, Harvard Medical School, Boston, Massachusetts 02115.

Page 4790, first column, 15th line from the bottom: Read "the molecular weight of *cyclo*(Pro-Gly)₁", rather than "the molecular weight of *cyclo*(Pro-Gly)₂".

Page 4797, first column, line 12: Change "Table IV" to read "Table II".

Page 4797, column 2, third paragraph, line 14: Change "Table VIII" to read "Table IV".

Page 4797, fourth paragraph, lines 11 and 12: Change "Figure 4 and Table X" to read "Figure 2 and Table XI".

Application of Linear Dichroism to the Analysis of Electronic Absorption Spectra of Biphenyl, Fluorene, 9,9'-Spirofluorene, and [6.6]Vespirene. Interpretation of the Circular Dichroism Spectrum of [6.6]Vespirene [*J. Am. Chem. Soc.*, **99**, 6861 (1977)]. By JACOB SAGIV,* AMNON YOGEV, and YEHUDA MAZUR, Department of Isotopes and Department of Organic Chemistry, The Weizmann Institute of Science, Rehovot, Israel.

Page 6868, column 1, line 1 should be: "... fluorene transitions lead to *z*- and *y*-polarized exciton pairs having opposite sign in the CD spectrum, while the transversal *x*-polarized bands ...".